Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

Diazidobis[4,4,5,5-tetramethyl-2-(1,3-thiazol-2-yl)-2-imidazoline-1-oxyl-3-oxide-κ^2O,N]manganese(II)

Jiu Li Chang, Zhi Yong Gao and Kai Jiang*

College of Chemistry and Environmental Science, Henan Normal University, Xinxiang 453002, People's Republic of China
Correspondence e-mail: gaozhy201@sohu.com

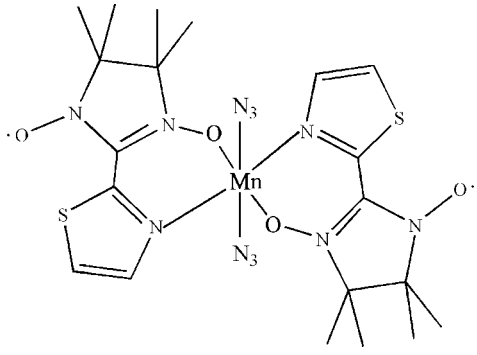
Received 25 December 2008; accepted 7 January 2009

Key indicators: single-crystal X-ray study; T = 291 K; mean σ(C–C) = 0.003 Å; R factor = 0.039; wR factor = 0.110; data-to-parameter ratio = 16.8.

In the crystal structure of the title compound, $[Mn(N_3)_2(C_{10}H_{14}N_3O_2S)_2]$, the Mn(II) atom exhibits a roughly octahedral coordination geometry. The Mn(II) atom lies on an inversion centre, thus the asymmetric unit comprises one half-molecule. The metal center is six-coordinated by two azide anions and by two chelating 4,4,5,5-tetramethyl-2-(1,3-thiazol-2-yl)-2-imidazoline-1-oxyl-3-oxide nitronyl nitroxide radical ligands, leading to two six-membered chelate rings.

Related literature

For the design and synthesis of molecule-based magnetic materials, see: Aoki *et al.* (2003). For nitronyl nitroxide radicals, see: Minguet *et al.* (2000); Catala *et al.* (2005). For transition metal–radical complexes, see: Wang *et al.* (2005). For paramagnetic metal complexes of nitronyl nitroxide radicals, see: Li *et al.* (2002); Liu *et al.* (2001). For the synthesis, see: Ullman *et al.* (1970, 1972)



Experimental

Crystal data

$[Mn(N_3)_2(C_{10}H_{14}N_3O_2S)_2]$
M_r = 619.60
Monoclinic, $P2_1/c$
a = 9.9600 (18) Å
b = 12.272 (2) Å
c = 11.353 (2) Å
β = 103.714 (3)°
V = 1348.1 (4) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.70 mm^{-1}
T = 291 (2) K
0.45 × 0.30 × 0.25 mm

Data collection

Bruker SMART CCD area-detector diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.745, T_{max} = 0.846
7966 measured reflections
3061 independent reflections
2628 reflections with $I > 2\sigma(I)$
R_{int} = 0.036

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.039
$wR(F^2)$ = 0.110
S = 1.06
3061 reflections
182 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.37 e Å^{-3}
$\Delta\rho_{min}$ = −0.25 e Å^{-3}

Data collection: *SMART* (Bruker, 2002); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *publCIF* (Westrip, 2009).

This work was supported by the National Natural Science Foundation of China (grant No. 20471026) and the Natural Science Foundation of Henan Province (grant No. 0311021200).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: KP2202).

References

Aoki, C., Ishida, T. & Nogami, T. (2003). *Inorg. Chem.* **42**, 7616–7625.
Bruker (2002). *SAINT and SMART*. Bruker AXS Inc., Madison, Winsonsin, USA.
Catala, L., Moigne, J. L., Gruber, N., Novoa, J. J., Rabu, P., Belorizky, E. & Turek, P. (2005). *Chem. Eur. J.* **11**, 2440–2454.
Li, L. C., Liao, D. Z., Jiang, Z. H. & &Yan, S. P. (2002). *J. Chem. Soc. Dalton Trans.* pp. 1350–1353.
Liu, Z. L., Zhao, Q. H., Li, S. Q., Liao, D. Z. & Jiang, Z. H. (2001). *Inorg. Chem. Commun.* **4**, 322–325.
Minguet, M., Amabilino, D. B., Cirujeda, J., Wurst, K., Mata, I., Molins, E., Novoa, J. J. & Veciana, J. (2000). *Chem. Eur. J.* **6**, 2350–2361.
Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Ullman, E. F., Call, L. & Osieckei, J. H. J. (1970). *J. Org. Chem.* **35**, 3623–3628.
Ullman, E. F., Osiecki, J. H., Boocock, D. G. B. & Darcy, R. (1972). *J. Am. Chem. Soc.* **94**, 7049–7059.
Wang, L.-Y., Chang, J.-L., Jiang, K., Ma, L.-F. & Wang, Y.-F. (2005). *Acta Cryst.* E**61**, m2230–m2231.
Westrip, S. P. (2009) *publCIF*. In preparation.

supplementary materials

Acta Cryst. (2009). E**65**, m181 [**doi:10.1107/S1600536809000786**]

Diazidobis[4,4,5,5-tetramethyl-2-(1,3-thiazol-2-yl)-2-imidazoline-1-oxyl-3-oxide-κ^2O,N]manganese(II)

J. L. Chang, Z. Y. Gao and K. Jiang

Comment

The design and synthesis of molecule-based magnetic materials is one of the major subjects of materials science(Aoki *et al.* 2003). In many different types of organic radicals, research has focused on the nitronyl nitroxide radicals (NITR) family because of their flexibility and functionality (Minguet *et al.* 2000; Catala *et al.* 2005). The nitroxide derivatives can be bound to the metal center through the oxygen atoms of O–N groups, affording a good variety of transition metal–radical complexes (Wang *et al.* 2005;). There have been many magnetic studies on transition metal complexes with nitronyl nitroxide and imino nitroxide radicals and paramagnetic metal complexes of nitronyl nitroxide radicals have been extensively studied (Li *et al.* 2002; Liu *et al.* 2001). In the present paper, we report the synthesis and crystal structure of the title compound $Mn(N_3)_2(NIT2\text{-}thz)_2$.

Experimental

NIT2-thz [NIT2-thz = 4,4,5,5-tetramethyl-2-(1,3-thiazol-2-yl)-2-imidazoline-1-oxyl-3-oxide] was synthesized using a method in the literatrue (Ullman *et al.* 1970; Ullman *et al.* 1972). Mn $(Ac)_2$. $4H_2O$(1 mmol) and NIT2-thz (2 mmol) were mixed in 30 ml of methanol. An aqueous solution (10 ml) of NaN_3 (2 mmol) was added to this solution. The mixture was stirred for an 1 h and filtered off. The filtrate was kept at room temperatrue for 1 meek, and well formed dark brown crstals of $Mn(N_3)_2(NIT2\text{-}thz)_2$ were obtained.

Refinement

The H atoms were positioned geometrically and refined using the riding-model approximation, with C—H = 0.93 or 0.96Å and U_{iso}(H) = 1.2Ueq(carrier) or U_{iso}(H) = 1.5Ueq(methyl carrier).

Figures

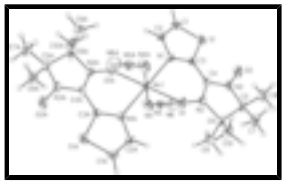


Fig. 1. *ORTEP* drawing of the title compound with atom labeling. The thermal ellipsoids are drawn at 30% probability level.[symmetry codes: *x*,-*y*,-*z* + 1].

Diazidobis[4,4,5,5-tetramethyl-2-(1,3-thiazol-2-yl)-2-imidazoline-1-oxyl- 3-oxide-κ^2O,N]manganese(II)

Crystal data

$[Mn(N_3)_2(C_{10}H_{14}N_3O_2S)_2]$	F_{000} = 642
M_r = 619.60	D_x = 1.526 Mg m^{-3}
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation λ = 0.71073 Å
Hall symbol: -P 2ybc	Cell parameters from 4318 reflections
a = 9.9600 (18) Å	θ = 2.5–28.2º
b = 12.272 (2) Å	μ = 0.70 mm^{-1}
c = 11.353 (2) Å	T = 291 (2) K
β = 103.714 (3)º	Block, dark brown
V = 1348.1 (4) Å^3	0.45 × 0.30 × 0.25 mm
Z = 2	

Data collection

Bruker SMART CCD area-detector diffractometer	3061 independent reflections
Radiation source: fine-focus sealed tube	2628 reflections with $I > 2\sigma(I)$
Monochromator: graphite	R_{int} = 0.036
T = 291(2) K	θ_{max} = 27.5º
φ and ω scans	θ_{min} = 2.5º
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	h = −12→12
T_{min} = 0.745, T_{max} = 0.846	k = −15→15
7966 measured reflections	l = −12→14

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)]$ = 0.039	H-atom parameters constrained
$wR(F^2)$ = 0.110	$w = 1/[\sigma^2(F_o^2) + (0.0605P)^2 + 0.2559P]$ where $P = (F_o^2 + 2F_c^2)/3$
S = 1.06	$(\Delta/\sigma)_{max} < 0.001$
3061 reflections	$\Delta\rho_{max}$ = 0.37 e Å^{-3}
182 parameters	$\Delta\rho_{min}$ = −0.25 e Å^{-3}
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes)

are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Mn1	0.0000	0.0000	0.5000	0.03345 (14)
S1	0.25159 (5)	0.27889 (4)	0.36683 (5)	0.05049 (17)
O1	0.17034 (15)	−0.08711 (11)	0.44955 (14)	0.0504 (4)
O2	0.38630 (18)	0.14599 (13)	0.23568 (17)	0.0660 (5)
N1	0.12384 (15)	0.14413 (11)	0.47278 (13)	0.0337 (3)
N2	0.23912 (14)	−0.05108 (11)	0.37542 (13)	0.0336 (3)
N3	0.34509 (15)	0.05765 (13)	0.27421 (14)	0.0407 (4)
N4	−0.1080 (2)	0.00819 (18)	0.31193 (18)	0.0641 (6)
N5	−0.08511 (17)	0.04681 (12)	0.22465 (15)	0.0433 (4)
N6	−0.0640 (3)	0.08223 (18)	0.13715 (19)	0.0761 (7)
C1	0.1569 (2)	0.32772 (15)	0.4616 (2)	0.0495 (5)
H1	0.1479	0.4012	0.4782	0.059*
C2	0.0969 (2)	0.24624 (14)	0.50955 (17)	0.0412 (4)
H2	0.0413	0.2584	0.5635	0.049*
C3	0.20718 (16)	0.14870 (12)	0.39806 (15)	0.0318 (3)
C4	0.25886 (16)	0.05337 (13)	0.34979 (15)	0.0317 (3)
C5	0.40415 (19)	−0.05207 (16)	0.25646 (16)	0.0406 (4)
C6	0.30045 (18)	−0.12873 (14)	0.29998 (16)	0.0382 (4)
C7	0.4089 (3)	−0.0651 (2)	0.12397 (19)	0.0628 (6)
H7A	0.3179	−0.0551	0.0732	0.094*
H7B	0.4417	−0.1367	0.1114	0.094*
H7C	0.4702	−0.0116	0.1039	0.094*
C8	0.5497 (2)	−0.0545 (2)	0.3366 (2)	0.0601 (6)
H8A	0.6029	0.0037	0.3137	0.090*

H8B	0.5923	−0.1230	0.3269	0.090*
H8C	0.5461	−0.0455	0.4197	0.090*
C9	0.1798 (2)	−0.1673 (2)	0.1994 (2)	0.0595 (6)
H9A	0.1102	−0.1990	0.2346	0.089*
H9B	0.2118	−0.2208	0.1507	0.089*
H9C	0.1413	−0.1064	0.1496	0.089*
C10	0.3643 (3)	−0.22455 (19)	0.3780 (2)	0.0620 (6)
H10A	0.4243	−0.1981	0.4515	0.093*
H10B	0.4167	−0.2680	0.3345	0.093*
H10C	0.2925	−0.2682	0.3974	0.093*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Mn1	0.0362 (2)	0.0333 (2)	0.0373 (2)	−0.00148 (13)	0.02166 (16)	0.00206 (13)
S1	0.0483 (3)	0.0342 (2)	0.0743 (4)	−0.00716 (19)	0.0251 (3)	0.0074 (2)
O1	0.0597 (9)	0.0351 (6)	0.0717 (10)	0.0065 (6)	0.0459 (8)	0.0095 (6)
O2	0.0697 (10)	0.0592 (9)	0.0861 (12)	−0.0037 (8)	0.0525 (10)	0.0169 (8)
N1	0.0365 (7)	0.0299 (7)	0.0381 (7)	0.0019 (5)	0.0156 (6)	−0.0003 (5)
N2	0.0325 (7)	0.0336 (7)	0.0395 (8)	0.0037 (5)	0.0182 (6)	0.0003 (6)
N3	0.0371 (8)	0.0466 (8)	0.0452 (8)	0.0000 (6)	0.0230 (7)	0.0038 (7)
N4	0.0681 (13)	0.0827 (14)	0.0416 (10)	−0.0251 (10)	0.0134 (9)	0.0060 (9)
N5	0.0497 (9)	0.0360 (8)	0.0453 (9)	−0.0042 (6)	0.0138 (7)	0.0000 (7)
N6	0.121 (2)	0.0587 (12)	0.0582 (12)	−0.0128 (12)	0.0400 (13)	0.0104 (10)
C1	0.0490 (11)	0.0305 (8)	0.0668 (13)	0.0003 (8)	0.0093 (10)	−0.0061 (8)
C2	0.0449 (10)	0.0352 (8)	0.0444 (10)	0.0065 (7)	0.0122 (8)	−0.0065 (7)
C3	0.0297 (8)	0.0302 (7)	0.0372 (8)	−0.0012 (6)	0.0113 (6)	0.0028 (6)
C4	0.0277 (8)	0.0361 (8)	0.0335 (8)	0.0006 (6)	0.0118 (6)	0.0029 (6)
C5	0.0346 (9)	0.0556 (11)	0.0361 (9)	0.0076 (7)	0.0170 (7)	−0.0030 (8)
C6	0.0369 (9)	0.0404 (9)	0.0402 (9)	0.0095 (7)	0.0151 (7)	−0.0048 (7)
C7	0.0640 (14)	0.0898 (17)	0.0422 (11)	0.0087 (12)	0.0277 (10)	−0.0064 (11)
C8	0.0358 (10)	0.0856 (16)	0.0599 (13)	0.0054 (10)	0.0133 (10)	−0.0127 (12)
C9	0.0569 (13)	0.0618 (13)	0.0587 (13)	−0.0084 (10)	0.0117 (11)	−0.0173 (10)
C10	0.0695 (15)	0.0542 (12)	0.0703 (15)	0.0292 (11)	0.0324 (13)	0.0111 (11)

Geometric parameters (Å, °)

Mn1—N4^{i}	2.153 (2)	C2—H2	0.9300
Mn1—N4	2.153 (2)	C3—C4	1.438 (2)
Mn1—O1^{i}	2.1931 (12)	C5—C8	1.518 (3)
Mn1—O1	2.1931 (12)	C5—C7	1.525 (3)
Mn1—N1	2.2194 (14)	C5—C6	1.561 (3)
Mn1—N1^{i}	2.2194 (14)	C6—C10	1.518 (3)
S1—C1	1.699 (2)	C6—C9	1.524 (3)
S1—C3	1.7170 (16)	C7—H7A	0.9600
O1—N2	1.2832 (17)	C7—H7B	0.9600
O2—N3	1.273 (2)	C7—H7C	0.9600
N1—C3	1.321 (2)	C8—H8A	0.9600

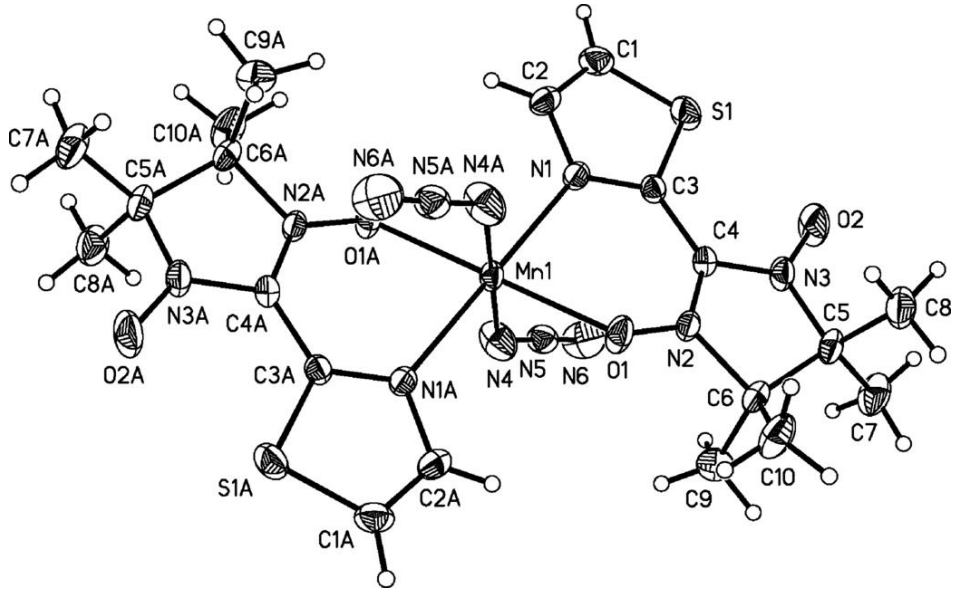
N1—C2	1.367 (2)	C8—H8B	0.9600
N2—C4	1.339 (2)	C8—H8C	0.9600
N2—C6	1.504 (2)	C9—H9A	0.9600
N3—C4	1.3513 (19)	C9—H9B	0.9600
N3—C5	1.502 (2)	C9—H9C	0.9600
N4—N5	1.168 (2)	C10—H10A	0.9600
N5—N6	1.148 (2)	C10—H10B	0.9600
C1—C2	1.345 (3)	C10—H10C	0.9600
C1—H1	0.9300		

$N4^{i}$—Mn1—N4	180.0	N2—C4—C3	127.70 (14)
$N4^{i}$—Mn1—$O1^{i}$	89.95 (8)	N3—C4—C3	123.31 (15)
N4—Mn1—$O1^{i}$	90.05 (8)	N3—C5—C8	106.60 (17)
$N4^{i}$—Mn1—O1	90.05 (8)	N3—C5—C7	109.34 (17)
N4—Mn1—O1	89.95 (8)	C8—C5—C7	109.88 (17)
$O1^{i}$—Mn1—O1	180.0	N3—C5—C6	100.84 (13)
$N4^{i}$—Mn1—N1	90.68 (6)	C8—C5—C6	114.10 (17)
N4—Mn1—N1	89.32 (6)	C7—C5—C6	115.27 (18)
$O1^{i}$—Mn1—N1	97.92 (5)	N2—C6—C10	109.20 (15)
O1—Mn1—N1	82.08 (5)	N2—C6—C9	105.59 (15)
$N4^{i}$—Mn1—$N1^{i}$	89.32 (6)	C10—C6—C9	110.10 (19)
N4—Mn1—$N1^{i}$	90.68 (6)	N2—C6—C5	100.77 (13)
$O1^{i}$—Mn1—$N1^{i}$	82.08 (5)	C10—C6—C5	115.73 (16)
O1—Mn1—$N1^{i}$	97.92 (5)	C9—C6—C5	114.43 (16)
N1—Mn1—$N1^{i}$	180.0	C5—C7—H7A	109.5
C1—S1—C3	89.35 (9)	C5—C7—H7B	109.5
N2—O1—Mn1	124.73 (10)	H7A—C7—H7B	109.5
C3—N1—C2	110.83 (14)	C5—C7—H7C	109.5
C3—N1—Mn1	125.29 (11)	H7A—C7—H7C	109.5
C2—N1—Mn1	122.15 (11)	H7B—C7—H7C	109.5
O1—N2—C4	126.95 (13)	C5—C8—H8A	109.5
O1—N2—C6	120.47 (13)	C5—C8—H8B	109.5
C4—N2—C6	112.47 (13)	H8A—C8—H8B	109.5
O2—N3—C4	123.82 (15)	C5—C8—H8C	109.5
O2—N3—C5	123.29 (14)	H8A—C8—H8C	109.5
C4—N3—C5	112.24 (14)	H8B—C8—H8C	109.5
N5—N4—Mn1	135.07 (17)	C6—C9—H9A	109.5
N6—N5—N4	178.1 (2)	C6—C9—H9B	109.5
C2—C1—S1	111.16 (14)	H9A—C9—H9B	109.5
C2—C1—H1	124.4	C6—C9—H9C	109.5
S1—C1—H1	124.4	H9A—C9—H9C	109.5
C1—C2—N1	114.81 (16)	H9B—C9—H9C	109.5
C1—C2—H2	122.6	C6—C10—H10A	109.5
N1—C2—H2	122.6	C6—C10—H10B	109.5
N1—C3—C4	123.11 (14)	H10A—C10—H10B	109.5
N1—C3—S1	113.83 (12)	C6—C10—H10C	109.5
C4—C3—S1	123.04 (12)	H10A—C10—H10C	109.5

N2—C4—N3	108.87 (14)	H10B—C10—H10C	109.5
N4^{i}—Mn1—O1—N2	−122.97 (15)	O1—N2—C4—N3	175.59 (17)
N4—Mn1—O1—N2	57.03 (15)	C6—N2—C4—N3	−8.19 (19)
O1^{i}—Mn1—O1—N2	105 (8)	O1—N2—C4—C3	−0.6 (3)
N1—Mn1—O1—N2	−32.28 (14)	C6—N2—C4—C3	175.65 (17)
N1^{i}—Mn1—O1—N2	147.72 (14)	O2—N3—C4—N2	−178.27 (17)
N4^{i}—Mn1—N1—C3	118.52 (15)	C5—N3—C4—N2	−7.24 (19)
N4—Mn1—N1—C3	−61.48 (15)	O2—N3—C4—C3	−1.9 (3)
O1^{i}—Mn1—N1—C3	−151.43 (14)	C5—N3—C4—C3	169.12 (16)
O1—Mn1—N1—C3	28.57 (14)	N1—C3—C4—N2	−3.7 (3)
N1^{i}—Mn1—N1—C3	24.5 (17)	S1—C3—C4—N2	174.59 (14)
N4^{i}—Mn1—N1—C2	−77.84 (15)	N1—C3—C4—N3	−179.39 (16)
N4—Mn1—N1—C2	102.16 (15)	S1—C3—C4—N3	−1.1 (2)
O1^{i}—Mn1—N1—C2	12.20 (15)	O2—N3—C5—C8	70.0 (2)
O1—Mn1—N1—C2	−167.80 (15)	C4—N3—C5—C8	−101.13 (18)
N1^{i}—Mn1—N1—C2	−171.9 (18)	O2—N3—C5—C7	−48.8 (2)
Mn1—O1—N2—C4	26.1 (2)	C4—N3—C5—C7	140.13 (18)
Mn1—O1—N2—C6	−149.81 (13)	O2—N3—C5—C6	−170.64 (18)
N4^{i}—Mn1—N4—N5	−71 (2)	C4—N3—C5—C6	18.27 (18)
O1^{i}—Mn1—N4—N5	118.9 (3)	O1—N2—C6—C10	−42.4 (2)
O1—Mn1—N4—N5	−61.1 (3)	C4—N2—C6—C10	141.14 (17)
N1—Mn1—N4—N5	21.0 (3)	O1—N2—C6—C9	76.0 (2)
N1^{i}—Mn1—N4—N5	−159.0 (3)	C4—N2—C6—C9	−100.50 (18)
Mn1—N4—N5—N6	137 (8)	O1—N2—C6—C5	−164.65 (15)
C3—S1—C1—C2	−0.76 (16)	C4—N2—C6—C5	18.85 (18)
S1—C1—C2—N1	−0.1 (2)	N3—C5—C6—N2	−20.40 (16)
C3—N1—C2—C1	1.2 (2)	C8—C5—C6—N2	93.45 (18)
Mn1—N1—C2—C1	−164.58 (14)	C7—C5—C6—N2	−138.01 (17)
C2—N1—C3—C4	176.71 (16)	N3—C5—C6—C10	−138.01 (17)
Mn1—N1—C3—C4	−18.1 (2)	C8—C5—C6—C10	−24.2 (2)
C2—N1—C3—S1	−1.76 (19)	C7—C5—C6—C10	104.4 (2)
Mn1—N1—C3—S1	163.46 (8)	N3—C5—C6—C9	92.35 (18)
C1—S1—C3—N1	1.47 (15)	C8—C5—C6—C9	−153.79 (17)
C1—S1—C3—C4	−177.00 (16)	C7—C5—C6—C9	−25.3 (2)

Symmetry codes: (i) −*x*, −*y*, −*z*+1.

Fig. 1


C9A
C1
C2
S1
C7A
C10A
C6A
N6A
N5A
N4A
N1
C3
C5A
N2A
O1A
C4
O2
C8A
Mn1
N3
C5
C8
N3A
C4A
N4
N5
N6
O1
N2
O2A
C3A
N1A
C6
C7
S1A
C2A
C10
C9
C1A